CHEMISTRY

SOME BASIC CONCEPTS OF CHEMISTRY

01 NATURE OF MATTER

02 LAWS OF CHEMICAL COMBINATIONS

03 DALTON'S ATOMIC THEORY

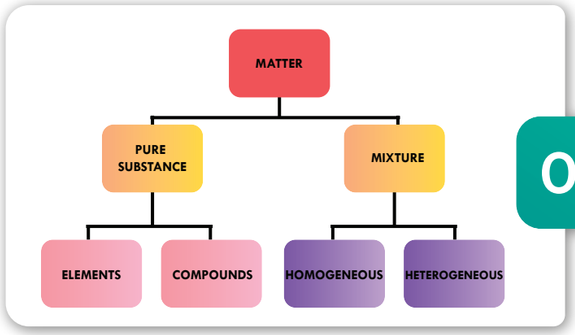
04 MOLE CONCEPT

05 PERCENTAGE COMPOSITION

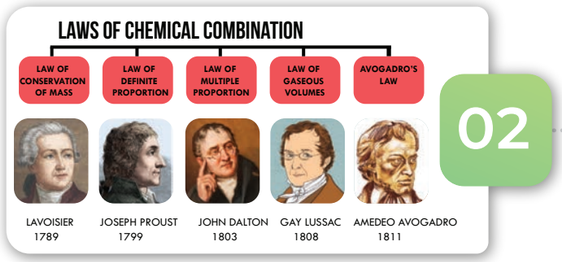
06 EF & MF

07 LIMITING REACTANT

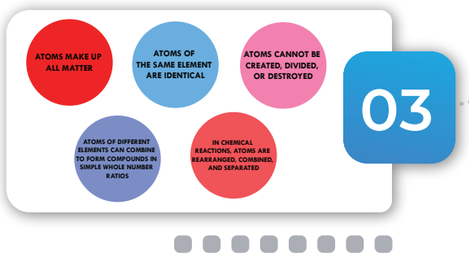
08 STOICHIOMETRIC CALCULATIONS

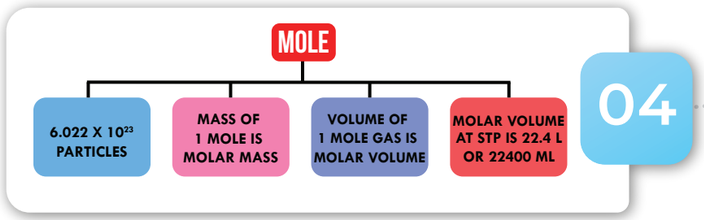


Q) Which one of the following is not a mixture ?
(A) Tap water (B) Distilled water
(C) Salt in water (D) Oil in water

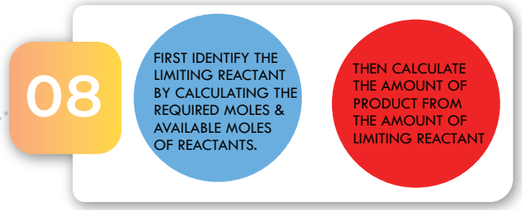


Q) Which one of the following pairs of compound illustrate the law of multiple proportions ?
(A) H_2O, Na_2O (B) MgO, Na_2O
(C) Na_2O, BaO (D) $SnCl_2$, $SnCl_4$

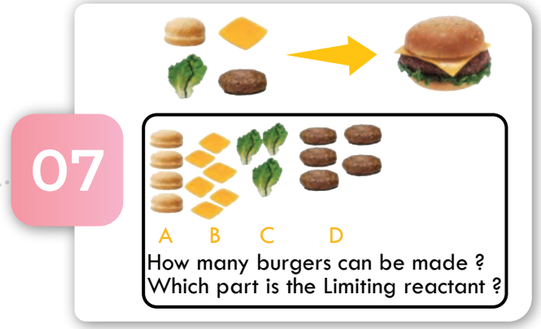




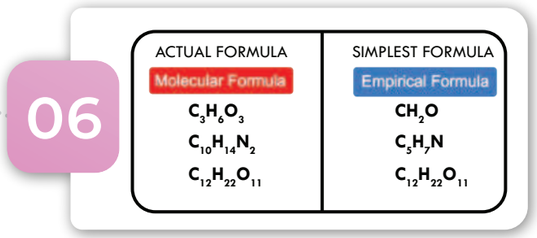
Q) Which one of the followings has maximum number of atoms ?
(A) 1 g of Mg(s) [Atomic mass of Mg = 24]
(B) 1 g of O_2 [Atomic mass of O=16]
(C) 1 g of Li(s) [Atomic mass of Li = 7]
(D) 1 g of Ag(s) [Atomic mass of Ag = 108]



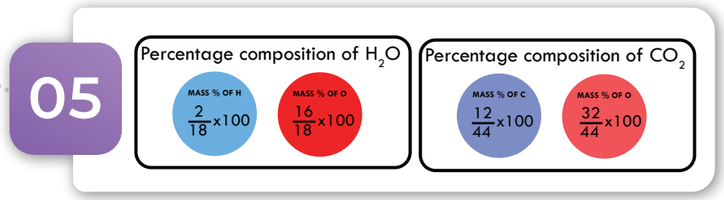
Q) When 22.4L of $H_2(g)$ is mixed with 11.2L of $Cl_2(g)$, each at STP, the moles of HCl(g) formed is equal to :
(A) 0.5 (B) 1.5 (C) 1 (D) 2



Q) The number of moles of hydrogen molecules required to produce 20 moles of ammonia through Haber's process is :
(A) 40 (B) 10 (C) 20 (D) 30



Q) An organic compound contains 80% (by wt.) C & the remaining percentage of H . The empirical formula of this compound is :
(A) CH_3 (B) CH_4 (C) CH (D) CH_2



Q) Mass % of carbon in ethanol is :
(A) 52 (B) 13 (C) 34 (D) 90